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Each problem includes the information relative to the problem itself, along with the necessary input data (the chemical symbols, molecules, atomic or molecular numbers, experimental conditions, etc.). The software displays each problem as a textual representation, ensuring easy understanding of the problem. Problems can be printed to pre-formatted text files, allowing to prepare personalized problem lists of any length and complexity. The software allows also to perform several calculations by textual commands. In particular the following can be done (by the simple choice of various possible operations): * Selection of primitive Elements, Compounds and molecular species of arbitrary size and composition, with graphical representation of their structures and list of their absolute compositions and experimental conditions * Selection and comparison of different experimental conditions (temperature, pressure, co-reactants, reactants, etc.) * Selection of possible reactions with graphical representation of the products and a brief representation of the reaction mechanism * Visualization of simple molecular structures, such as Molecules, Polymers, Carbohydrates, Nucleic Acids, Fatty Acids and Glycans * Calculation of molecular properties and comparison with a list of common experimental conditions * Calculation of molar masses, number of moles, net mass balances, number of moles of reactants, products and co-reactants * Calculation of mole fractions, mass fractions, molar fractions, charge fractions, numbers of charges, charge densities, atomic fractions and bonding properties (densities of states) * Calculation of Gibbs Energy and Equilibrium constants * Calculation of the extent of isomerism (amount of possible structural isomers) * Calculation of the electric charges of ions and molecules * Calculation of the ionic radii of ions and determination of the packing factor of ions * Calculation of the ionic dissociation constants * Calculation of electronegativity and chemical hardness of substances * Calculation of the adiabatic ionization potentials and chemical potentials of molecular species * Calculation of Henry's law and other empirical law constants of solubility (extraction, precipitation, adsorption, diffusion) * Calculation of diffusion rate constants in organic and inorganic solutions, including a graphical representation of the ions concentration profile in the presence of a gradient of external electric field and a schematic representation of the ions concentration profile in the presence of a diffusive gradient * Calculation of forward and reverse rate constants of chemical reactions * Calculation of the

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The software generates instantaneously and continuously new problems on all common topics of a basic Chemistry course, showing their solutions when prompted ; also features various useful Tools (Chem Calculator, Plotter, Balancer etc.). ## A.1.1 Compatibility: **All** versions of the software are compatible. **On Windows** : **xchem_ver9.1.zip** and **xchem_ver9.1.zip(32bit, 64bit)** - compatible versions with the 32 bits and 64 bits versions of the programming language C++ (windows version) **xchem_ver9.2.zip** and **xchem_ver9.2.zip(32bit, 64bit)** - compatible versions with the 32 bits and 64 bits versions of the programming language C (mac version) **xchem_ver9.3.zip** and **xchem_ver9.3.zip(32bit, 64bit)** - compatible versions with the 32 bits and 64 bits versions of the programming language C (linux version) **On Ubuntu 16.04** **xchem_ver9.1.deb** - compatible version with the 32 bits version of the debian software repository (Ubuntu) **xchem_ver9.2.deb** - compatible version with the 32 bits version of the debian software repository (Ubuntu) ## A.1.2 License: **Freedom Software, Inc.** **Estate of inventor: inventor's family** **Legal entity: U.S.A** ## A.1.3 How to install: On Windows or Ubuntu Linux (using a Debian base): 1) Install the 32 bits version of

the software: *xchem_ver9.1.zip(32bit)*, *xchem_ver9.1.zip(32bit)* 2) Extract the file *xchem_ver9.1.bat* in a folder of your choice. 3) Double-click on 09e8f5149f

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Each problem is composed of several steps of N text frames, each frame being a new situation, governed by an equation. Each step has an average complexity N, to which can be added an individual complexity (from 1 to N) and a personal complexity (random). The challenge of a problem is to find the amount of elemental matter required to produce a definite amount of a product of interest, from elementary reaction equations and the addition of stoichiometric constants, randomised values (from 1 to N) and various unpredictable features (from statistics to chemical jargon, sometimes based on common sense) - or even from fractions, percentages, operators.... These values must be computed for each step of the reaction, and passed in the text frames of the problem. While computing, the software can produce an abundance curve, showing the expected trend of the problem, and, if the required information is missing, it can also create it based on the individual complexity given in the problem text. The random amount, as well as the stoichiometric coefficients, can be freely selected between the upper and lower limits of the given range. The personal complexity can also be freely selected within the chosen range. The solution of a problem is based on entering the number of moles of each component and the sum of the stoichiometric coefficients, all expressed in moles, i.e. moles = grams. Only then can the software access the problem text of the problem, and calculate all the terms of the equation. It also re-estimates any missing stoichiometric coefficient, whether random or personal, and can produce a plot of the relevant graph of the problem. Once all components and stoichiometric coefficients have been entered, the software computes the relevant reaction terms (typically products, reactants, stoichiometric coefficients, partial reaction coefficients, absolute reaction coefficients, work, heat and entropy terms, enthalpy and Gibbs free energy terms...) and checks that only positive values (and not negative values) have been obtained. The software also makes sure the stoichiometric coefficients in use are within the observed range and checks the validity of the reaction based on the chemical information required. Once all reactions terms have been calculated, the software sums them up, repeating the summing until the problem has been completed, or until a predetermined length of time has been reached. Then it computes the number of moles of the products of interest, as well as all the terms of the Gibbs equation. Finally, the software checks the quantities actually used in

What's New in the Chemistry Problems?

======= ================ Molecular Model of Reaction Yield. o "Yield" is the fraction of the reactants that are converted to the desired product. o For a stoichiometric reaction, the reactants are present in the desired product in the same relative amounts; thus yield is equal to 100%. o For a stoichiometric reaction, the stoichiometric coefficients of the reactants (or products) must be 100%; thus yield is 100% in both cases. o In a pseudo-stoichiometric reaction, the stoichiometric coefficients do not add up to 100%; thus the yield is not necessarily 100%. o In a pseudo-stoichiometric reaction, the yield may be 100% or 0%. o In an endo-reaction, the yield is usually 0% because either reactants, products or both are degraded. o A variant of the stoichiometric reaction is the catalytic reaction, in which one of the reactants is consumed or produced in a different amount than the stoichiometric; in this case, the yield is usually between 0% and 100%. o A catalytic reaction may also have an exo-reaction, in which one or both reactants are produced. o Sometimes, the reactions are assumed to occur reversibly, and yield is then equal to the difference of the starting and final concentrations of reactants and products. o Similarly, yield may be equal to the sum of starting and final concentrations of reactants and products. o For a stepwise reaction, yield is only equal to the product of its initial and

final concentrations. o In a pseudo-stepwise reaction, yield is not necessarily the product of the concentrations at each step. o In a pseudo-stepwise reaction, yield may be equal to any value between the initial and final concentrations of reactants, products or both. o For a nucleophilic substitution, yield is equal to the relative amount of the desired product (over all possible products). o In a nucleophilic substitution, yield is usually 0% because usually the desired product is unreactive; it is formed in the same way as an intermediate of a reaction. o For a nucleophilic addition, yield is equal to the relative amount of the desired product. o In a nucleophilic addition, yield may be equal to the relative amounts of the starting materials. o In a nucleophilic addition, yield is usually 0% because usually

System Requirements For Chemistry Problems:

The game requires 1GB of RAM and 200MB of available space on your HDD. It can also be played in offline mode. Description: Don't be mistaken. This is not just another shooter. You are about to embark on an adventure and never will you return. Prepare yourself for the "Majestic of the Outlands" challenge, with your own handcrafted rifle. Get ready to explore a new universe, forge friendships with a wide range of characters, face numerous deadly dangers and test your ability to fight both human and alien soldiers. You are the last man standing

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